An ionic liquid process for mercury removal from natural gas


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Efficient scrubbing of mercury vapour from natural gas streams has been demonstrated both in the laboratory and on industrial scale, using chlorocuprate(II) ionic liquids impregnated on high surface area porous solid supports, resulting in the effective removal of mercury vapour from gas streams. This material has been commercialised for use within the petroleum gas production industry, and has currently been running continuously for three years on a natural gas plant in Malaysia. Here we report on the chemistry underlying this process, and demonstrate the transfer of this technology from gram to ton scale.

Introduction

Mercury is a natural component of the earth’s crust, and through a number of natural and anthropogenic cycles is released into the environment as a toxic pollutant. All forms of mercury viz. elemental mercury and oxidised, \( \text{Hg}(I) \) and \( \text{Hg}(II) \) are intrinsically toxic. One of the main sources of environmental pollution by mercury is as releases from fossil fuels, either during combustion of coal or emitted from natural gas. Many natural gas fields contain mercury released from mercury-containing ores through secondary geothermal processes. Natural gas fields typically contain mercury, in either elemental or combinations of elemental and organometallic forms, in concentrations in the range \(<0.1\text{-}5000 \mu \text{g m}^{-3}\) depending on geographical location and geology. Although these concentrations are usually relatively low, for a typical gas processing plant treating 250 MMSCFD (Million standard cubic feet per day \(=28300 \text{ m}^3 \text{ d}^{-1}\)) of gas, the cumulative quantities can be significant and can lead to problems through accumulation via condensation and amalgamation. High mercury concentrations in oil and gas production correlate with regions of high mercury emission to the environment (see Figure 1).

Mercury can be extremely corrosive, causing destructive damage to process equipment, particularly aluminium heat exchangers, through liquid metal embrittlement. For example, an explosion in 1973 at the Skikda liquefied natural gas plant in Algeria led to 27 fatalities and financial losses of \$1 \text{ billion} due to catastrophic failure of an aluminium heat exchanger through reaction with mercury contaminants.

Mercury control and treatment within the oil and gas supply chain is increasingly recognised as important, both to protect equipment and personnel from deleterious exposure, and in order to comply with increasingly stringent discharge regulations. For natural gas processing, mercury control usually signifies the use of fixed-bed scrubbers containing a solid absorbent that can capture mercury vapour either through absorption, amalgamation, or oxidation followed by absorption.
The most common approaches employ commercially available scrubbers, as shown in Table 1, incorporating sulfur, metal oxides/sulfides, or metals (particularly silver) as active agents on porous alumina, zeolite, or activated carbon supports.7

Table 1 Various mercury removal systems for natural gas streams.7

<table>
<thead>
<tr>
<th>Active Compound</th>
<th>Support/medium</th>
<th>Fate of mercury</th>
</tr>
</thead>
<tbody>
<tr>
<td>Sulfur</td>
<td>Carbon/Alumina</td>
<td>HgS</td>
</tr>
<tr>
<td>Metal sulfide</td>
<td>Carbon/Alumina</td>
<td>HgS</td>
</tr>
<tr>
<td>Silver</td>
<td>Zeolite</td>
<td>Ag-Hg amalgam</td>
</tr>
<tr>
<td>Thiol/Oxidising</td>
<td>Scavenger</td>
<td>Soluble Hg(II)</td>
</tr>
<tr>
<td>agent/chelating agent</td>
<td>solution</td>
<td>compound</td>
</tr>
<tr>
<td>Metal oxide/sulfide</td>
<td>Metal Oxide</td>
<td>HgO/HgS</td>
</tr>
</tbody>
</table>

Sulfur-impregnated activated carbons are probably the most widely used adsorbents for mercury control within the gas industry. Impregnation of sulfur into activated carbons overcomes some of the slow kinetics of the reaction of mercury with elemental sulfur under ambient conditions. Vidić and co-workers have proposed that this is may be due to the involvement of sulfur allotropes with greater number reactive end groups.8 However, activated carbons are not as mechanically robust as oxide supports such as alumina or silica, and therefore suffer from attrition which shortens their operational life-times.7 Mercury capture from gas using copper(II) chloride impregnated on carbon as a ‘chloride’ source has been investigated,9 however the resulting mercury(II) chloride was incompletely captured and leached under high pressure conditions.10

Ionic liquids11 have also been explored for liquid/liquid partitioning of mercury, as mercury(II), from water,12 initially using hydrophobic ionic liquids containing pendant sulfur ligands.13 Subsequently, it was shown that mercury(II) can partition into unfuctionalised hydrophobic ionic liquids.14,15

The effective inviolability of most ionic liquids means that gas-liquid contacting scenarios can be considered without contaminating process gas streams. Pinto and co-workers have reported the use of ionic liquids to capture elemental and oxidised mercury from combustion flue gases, combining the liquid with permanganate(VII) as an oxidant.16 Rogers and Holbrey17 have suggested using ionic liquids containing perhalide anions for the oxidative dissolution of metals, and Sasson and co-workers18 have reported using similar ionic liquid systems as liquid scrubbers to remove mercury from power-plant combustion gas. However, the corrosive nature of these highly oxidising ionic liquids (analogous to chlorine or bromine) to metals such as iron may place restrictions on their applicability.

Building on previous work developing ionic liquid approaches to oxidative dissolution,19,20 and selective extraction and separations,21 we considered whether ionic liquids22 incorporating metal complexes might oxidise elemental mercury, leading to the formation of stable anionic mercurate(II) species. Such anions should then become integral components of a new, more complex, ionic liquid system,23 which would also be non-volatile. Here, we report on the investigation of chlorocuprate(II)-based ionic liquids for the direct oxidation of elemental mercury,24 and the use of these ionic liquids in a solid-supported ionic liquid phase (SILP)25 for reactive capture of mercury from gas streams, which has led to the scale-up and deployment on industrial plants in Malaysia.26

Experimental

\(^{199}\)Hg NMR spectra were measured using a Bruker 500DRX spectrometer with a 1M solution of HgCl₂ (Sigma Aldrich) in dmso-d₆ as an internal standard (-1501 ppm relative to Hg(CH₃)₂) and deuterium lock.27 IR spectroscopy was carried out using a Perkin Elmer Spectrum 100 FT-IR spectrophotometer with a Universal Attenuated Total Reflectance (ATR) diamond top plate.

Mercury solubility in ionic liquids was measured by dissolving an accurately weighed sample of the ionic liquid phase (ca. 0.05 g) in water (10 cm³), removing the small quantity of precipitate produced by filtration, and then diluting to 50 cm³, followed by subsequent dilution of a 1 cm³ aliquot to 50 cm³. A sample (<50 μL) calculated to contain less than 1000 ng mercury was then analysed for mercury using a Milestone DMA-80 direct mercury analyser. Based on the sample size, initial mass of ionic liquid used for the contact test, and mercury content determined, the wt% of water soluble mercury present in the ionic liquid was calculated.

Mercury capture from the carrier gas by SILPs was tested using an accelerated breakthrough test rig constructed in-house, incorporating a Sir Galahad mercury generator and analyser (PS Analytical). SILPs (and commercial activated carbons for comparison) were crushed and sieved using a 300-500 μm mesh. Adsorbents (30-100 mg) were packed into a thermostatted glass column of diameter 1 mm. Mercury removal from test gas streams (nitrogen and methane) was examined using a high flow rate (600 cm³ min⁻¹) and high mercury content (2000 ng 1⁻¹) in the carrier gas; the outlet mercury concentration was monitored continuously.

Crystallography: A suitable crystal of [N₄⁺][Cu₂Cl₄] was selected and measured on a Rigaku Saturn724+ (2x2 bin mode) diffractometer. The crystal was kept at 120.0 K during data collection. Using Olex2,28 the structure was solved with the ShelixXS29 structure solution program using Direct Methods and refined with the ShelXL refinement package using Least Squares minimisation. Crystal Data for C₁₃H₁₆Cl₄CuN (M = 370.27 g mol⁻¹): monoclinic, space group P2₁/n (no. 14), \(a = 9.2771(4) \text{ Å}, b = 15.6784(5) \text{ Å}, c = 12.5521(5) \text{ Å}, \beta = 100.016(2)°, V = 1797.88(12) \text{ Å}^³, Z = 4, T = 120.0 \text{ K}, \mu(\text{MoKα}) = 1.647 mm⁻¹, D_{calc} = 1.368 g cm⁻³, 21158 reflections measured (5.98° ≤ 2θ ≤ 54.96°), 4122 unique (R_{int} = 0.0469, R_{wp} = 0.0382) which were used in all calculations. The final R₁ was 0.0377 (>2σ(I)) and wR₂ was 0.0951 (all data).

Results and discussion

Synthesis and characterisation of copper(II) ionic liquids

Chlorocuprate(II) ionic liquids were prepared by combining 1-butyl-3-methylimidazolium chloride ([C₄mim]Cl), 1-ethyl-3-
methylimidazolium chloride ([C₄mim]Cl), or tributylmethylammonium chloride ([N₄ 4 4 i]Cl, (ex. Sigma Aldrich) with copper(II) chloride (either anhydrous or dihydrate, ex. Sigma Aldrich) in methanol in either a 1:1 or 2:1 molar ratio, followed by solvent removal under reduced pressure, and then in vacuo at 80 °C. It is convenient to define the ionic liquid systems by the mole fraction χ_{CuCl₂}, which is defined by χ_{CuCl₂} = n_i/(n_i+n_e), where n_i is the number of moles of CuCl₂ and n_e is the number of moles of cation. The resulting ionic liquids formed as dark yellow-brown liquids which, in all cases, slowly solidified on standing at room temperature. Additionally, a 1:2:1 choline chloride/EG/CuCl₂·2H₂O (where EG = diethylene glycol) deep eutectic was prepared as a dark green liquid. The ionic liquids were characterised by DSC, TGA, microanalysis, electronic absorption and vibrational spectroscopy, Karl Fischer titration and surface analysis by SEM/EDAX.

Small single crystals were isolated from the bulk solidified [N₄ 4 4 i]CuCl₄ ionic liquid (1:1 CuCl₂:[N₄ 4 4 i]Cl), and analysed by X-ray crystallography. The structure confirmed the presence of a dimeric copper(II) dianion with four-coordinate copper centres separated by 3.33 Å (Figure 2). Li et al.31 have reported that [C₄mim]Cu/copper(II) chloride mixtures form [CuCl₃]²⁻ chlorocuprate(II) anions even in the presence of substantial concentrations of water. It was thus not anticipated that copper speciation would differ in ionic liquids formed with anhydrous or hydrated copper(II) chloride. Crystal structures of many chlorocuprate(II) salts with a variety of organic cations have been previously reported,32 and are dominated by salts of the discrete tetrachlorocuprate(II) anion, which can vary in structure from tetrahedral to square planar, with many intermediate geometries.

Reaction of copper(II) ionic liquids with mercury

Because the dominant form of mercury in natural gas is elemental, tests were conducted to determine the solubility of Hg(0) in the ionic liquids by heating a droplet of mercury (Sigma Aldrich) with the ionic liquids at 60 °C overnight in a sealed vial. No precautions to exclude air were made during sample loading. In each case, a colour change was observed in the ionic liquid, with the characteristic brown colour of the chlorocuprate(II)-based systems changing to (i) a clear pale green ionic liquid at χ_{CuCl₂} = 0.33, or to (ii) a pale orange-yellow ionic liquid with precipitation of a pale grey-white solid at χ_{CuCl₂} = 0.50, as shown in Figure 3. After heating, the size of the mercury drops remaining in the vials had visibly reduced and the mercury content in the ionic liquid phase was determined by mercury analysis. The results are shown in Table 2.

\[
\text{Hg}(0) + 2\text{Cu(II)} \rightarrow \text{Hg(II)} + 2\text{Cu(I)}
\]

Over 15 wt% mercury dissolution was observed (Table 2) for each of the chlorocuprate(II) ionic liquids tested, as determined from the water soluble mercury content of diluted aliquots of the ionic liquids. Reactive dissolution with similar end results was observed for the ionic liquids prepared from anhydrous and hydrated copper(II) chloride. A similar ability to dissolve mercury was observed for both χ_{CuCl₂} = 0.33 and 0.50, although the apparent outcomes of the dissolution process appear to be slightly different. The amount of mercury that was solubilised by the ionic liquids can be expressed by the molar concentration ratio [Cu(II)ₙ⁺]:[Hg(II)ₙ⁻] (in Table 2) from the initial concentration of copper in the ionic liquids and mass of mercury extracted. In all cases, a small excess of unreacted copper(II) in solution was apparent due to the residual paler brown colouration. This would imply that one mole of mercury was required to reduce two moles of copper.
The simplest mechanisms that can be proposed for this reaction, depending on the stoichiometry of the chlorocuprate(II) ionic liquids, are shown in Equations (1) to (3).

\[
\begin{align*}
2[CuCl]\text{\textsuperscript{2+}} + Hg(0) & \rightarrow [HgCl]\text{\textsuperscript{2+}} + 2[CuCl]^- \quad (1) \\
2[CuCl]\text{\textsuperscript{2+}} + 2Hg(0) & \rightarrow [Hg_2Cl]\text{\textsuperscript{2+}} + 2[CuCl]^- + 2CuCl(s) \quad (2) \\
[CuCl]\text{\textsuperscript{2+}} + Hg(0) & \rightarrow HgCl_2(s) + 2[CuCl]^- \quad (3)
\end{align*}
\]

The changes in colour of the copper(II)-containing ionic liquids after reaction with mercury are consistent with reduction to copper(I). Bolkan and Yoke described copper(I) systems ([C\text{\textsuperscript{2}}\text{\textsuperscript{3} mim\text{\textsuperscript{1}}}]CuCl\text{\textsuperscript{2}}; 0.33 < \chi_{CuCl} < 0.67) as having colours that varied from pale yellow to light green with increasing \chi_{CuCl}, and that darken rapidly in air and become paramagnetic as copper(II) is formed.

In Equation (1), all the products are anionic, and will incorporate into the ionic liquid. Equations (2) and (3) represent two contrasting interpretations of the observed reaction. In equation (2), copper(I) chloride precipitates and in equation (3), mercury(II) chloride precipitates. In the former case, the remaining solution contains both mercury(II) and copper(I) species; in the latter case, there will be no mercury in solution.

In Equation (2), it is speculated that the speciation of mercury is [Hg\text{\textsuperscript{2}}Cl\text{\textsuperscript{2+}}]. There is no evidence that this would be the speciation as it may be a monomeric species\(^{35}\) or a polymeric species,\(^{35}\) but the balance of probability is that it is likely to be dimeric.\(^{36}\) Whatever the speciation, it does not affect the validity of the following arguments.

Examples of \(^{199}\text{Hg}\) NMR spectra taken from \{[C\text{\textsuperscript{4} mim\text{\textsuperscript{1}}}]Cl-CuCl\text{\textsuperscript{2}} (\chi_{CuCl} = 0.50) + Hg(0)\} and \{[C\text{\textsuperscript{4} mim\text{\textsuperscript{1}}}]Cl-CuCl\text{\textsuperscript{2}} (\chi_{CuCl} = 0.33) + Hg(0)\} systems are shown in Figure 4. Well defined \(^{199}\text{Hg}\) NMR signals were obtained in each case, but with different chemical shifts. For \chi_{CuCl} = 0.33, this peak was found at -1106 ppm (relative to 1M HgCl\text{\textsuperscript{2}} in dmso); for \chi_{CuCl} = 0.50, the peak appeared at -1140 ppm. These positions are consistent with the formation of chloromercurate(II) anions,\(^{37}\) where the mercury is in the environment of four chlorine atoms. The lower field peak is assigned to [HgCl]\text{\textsuperscript{2-}} and the higher peak to [Hg\text{\textsuperscript{2}}Cl\text{\textsuperscript{2-}}]. As it is known that [HgCl\text{\textsuperscript{2}}] has a peak at -1197 ppm,\(^{30}\) this suggests that the mercury complex anion formed is dimeric rather than monomeric. This provides strong support for the reaction schemes outlined in Equations (1) and (2). Equation (3) has been eliminated as it would imply the absence of mercury in solution. The formation of a precipitate from the reaction of liquid mercury with the [C\text{\textsuperscript{4} mim\text{\textsuperscript{1}}}]CuCl\text{\textsuperscript{2}} (\chi_{CuCl} = 0.50) ionic liquid combined with the presence of chloromercurate(II) anions in solution as shown from \(^{199}\text{Hg}\) NMR spectroscopy supports the reaction mechanism in Equation (2) at this composition and the precipitation of CuCl as the reaction shifts the ionic liquid towards chloride deficient, \textit{i.e.} acidic, compositions. While little is known about the phase behaviour of chloromercurate(II) ionic liquids with \chi_{HgCl\text{\textsuperscript{2}}},\(^{18,39}\) the limiting point for homogeneity in chlorocuprate(II) systems is \chi_{CuCl} = 0.67,\(^{34}\) that is, with the copper:chloride ratio of 2:3.

### Copper(II) supported ionic liquids in the laboratory

The above results from the reactions of the copper(II) ionic liquids with elemental mercury give a good indication that they might be the basis of an absorption scrubber. This would combine the thermodynamically favourable oxidation of mercury(0) to mercury(II) by the copper(II), followed by complexation (and hence stabilisation) by the liberated chloride ions. Although in the laboratory, it is clearly possible to develop a mercury scrubber based on an ionic liquid system in the fluid state, the efficiency of passing natural gas through a viscous liquid system would prevent its scale-up for an industrial process. The logical way forward was to coat a dispersed solid with a selected and optimised ionic liquid to form a class of materials now recognised as a supported ionic liquid phase (SILP).\(^{25}\) This would lead to good contact between the gas and the ionic liquid, high throughput flow, and mechanical strength. A typical SILP is shown in Figure 5.

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Table 2 Solubility of bulk elemental mercury in metal-containing ionic liquid system at 60 °C (averages of three measurements).

<table>
<thead>
<tr>
<th>Ionic liquid composition</th>
<th>wt% mercury dissolved in IL</th>
<th>[Cu(II)\text{\textsuperscript{2+}}]:[Hg(II)\text{\textsuperscript{2+}}]</th>
</tr>
</thead>
<tbody>
<tr>
<td>[C\text{\textsuperscript{4} mim\text{\textsuperscript{1}}}]Cl-CuCl\text{\textsuperscript{2}}:2H\text{\textsuperscript{2}}O (1:1)</td>
<td>22.1 ± 1.4</td>
<td>2.6</td>
</tr>
<tr>
<td>[C\text{\textsuperscript{4} mim\text{\textsuperscript{1}}}]Cl-CuCl\text{\textsuperscript{2}} (2:1)</td>
<td>21.3 ± 2.3</td>
<td>2.2</td>
</tr>
<tr>
<td>[N\text{\textsuperscript{4} ethyl\text{\textsuperscript{1}}}]Cl-CuCl\text{\textsuperscript{2}} (2:1)</td>
<td>15.9 ± 0.3</td>
<td>2.4</td>
</tr>
<tr>
<td>[N\text{\textsuperscript{4} ethyl\text{\textsuperscript{1}}}]Cl-CuCl\text{\textsuperscript{2}}:2H\text{\textsuperscript{2}}O (2:1)</td>
<td>15.2 ± 0.1</td>
<td>2.3</td>
</tr>
<tr>
<td>[Chol]\text{\textsuperscript{1}}</td>
<td>Cl-EG-CuCl\text{\textsuperscript{2}}:2H\text{\textsuperscript{2}}O (1:2:1)</td>
<td>16.9 ± 0.1</td>
</tr>
</tbody>
</table>

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Fig. 4 \(^{199}\text{Hg}\) NMR (23 °C, neat, 89.57 MHz) spectra of the ionic liquids obtained from the reaction mixtures of (a) [C\text{\textsuperscript{4} mim\text{\textsuperscript{1}}}]Cl-CuCl\text{\textsuperscript{2}} + Hg(0) (red trace) and (b) [C\text{\textsuperscript{4} mim\text{\textsuperscript{1}}}]Cl-CuCl\text{\textsuperscript{2}} + Hg(0) (blue trace). The peak at -1501 ppm is the reference peak.

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Further, most mercury removal units used in the oil and gas industry are fixed-bed scrubbers, and so a SILP approach to heterogenise the ionic liquids would be a natural fit with standard industrial practice. SILPs, which incorporate high surface area thin-films of ionic liquids within a porous solid scaffold, have many advantages for efficient gas-liquid contacting when a static liquid phase (i.e. as a catalyst or adsorbent) is required.\textsuperscript{25}

SILPs containing the chlorocuprate(II) ionic liquids at 10 wt\% loading were prepared by the incipient wetness method, adding solutions of the respective ionic liquids in methanol to porous silica beads, followed by drying overnight at 80 °C. After drying, the SILPs were orange-yellow (see Figure 5), in contrast to the characteristic green of the methanolic solutions of the copper salts. ATIR spectra of SILPs showed no vibrational bands that could be assigned to methanol or to water, other than that residual on the silica surface (Figure 6).

Mercury adsorption from gas streams was tested by passing a mercury-containing gas (either N\textsubscript{2} or CH\textsubscript{4}) through a fixed bed of adsorbent, and comparing the inlet and outlet mercury concentrations. The more efficient the adsorbent is at removing mercury from the gas, the longer the time it takes for mercury to be detected at the outlet. The time from injection to detection is defined as the breakthrough time. Over the course of the measurements, the inlet mercury concentration was held at the high level of 2000 ng l\textsuperscript{-1} in order to accelerate the breakthrough tests. In packed adsorbent beds with no channeling, the outlet mercury concentration follows the ideal (and observed) profile illustrated in Figure 7. The outlet concentration remained constant at less than 0.2 ng l\textsuperscript{-1} prior to breakthrough, which was characterised by a rapid increase in its concentration. The breakthrough time was defined as the time at which mercury capture from the gas stream decreased to below 99.5% capture efficiency ([Hg]\textsubscript{out} > 10 ng l\textsuperscript{-1}). Characteristic mercury outlet profiles obtained under test conditions for a sulfur-impregnated activated carbon (ca. 10 wt\% sulfur) and for a chlorocuprate(II) containing SILP (10 wt\% [Na\textsubscript{4}+4\textsuperscript{+}]Cl-CuCl\textsubscript{2}, χ\textsubscript{CuCl\textsubscript{2}} = 0.50 on porous silica, 135 m\textsuperscript{2} g\textsuperscript{-1} surface area and 0.83 cm\textsuperscript{3} g\textsuperscript{-1} pore volume) are shown in Figure 7.

Under the conditions defined in Figure 7, the selected SILP remained active for 35 h until breakthrough, which corresponded to a mercury uptake of ca. 2.5 wt\% (2.52 mg, 1.27×10\textsuperscript{-5} mol). This 100 mg SILP sample contained 2.71×10\textsuperscript{-5} moles of copper(II), and so the breakthrough point represents 93% mercury capture efficiency based on the proposed capture mechanism requiring oxidation by two copper(II) centres. The activated carbon, in contrast, showed poorer performance, with a breakthrough of 12.7 h, which is equivalent to 9.0 wt\% mercury capture on the support. This was lower than anticipated, although this may be due to the extremely short residence times in these accelerated screening tests. Thus, in a laboratory test rig, the SILP outperformed the activated carbon
by a factor of 3, which was very encouraging for moving from laboratory to pilot plant.

Ionic liquids containing different cations ([Na(4,4-iPr)4]⁺, [C₆mim]⁺, and [C₄mim]⁺) were tested, resulting in similar breakthrough performance which appears to correlate with the total amount copper(II) within the tested samples (Figure 8). At higher loadings, a reduction in mercury capture performance was observed, presumably through plugging of the support through pore filling. This could be partially alleviated by changing to supports with large pore geometries. With theoretical maximum mercury capacities based on the mechanisms detailed in Equations (1) and (2). Following on from these small-scale laboratory tests, larger batches (ca. 100 g) of chlorocuprate(II) SILPs were prepared and tested in a 200 cm³ pilot-scale slipstream scrubber on-site at a Gas Processing Plant in Malaysia. The pilot trials (over 60 days) yielded a gas stream with mercury outlet concentrations of <0.01 µg m⁻³ (Figure 10), at least one order of magnitude below the sales specification for the natural gas.

After exhaustive extraction in contact with mercury-containing gas streams, the SILPs change colour from orange to pale green. This transformation is comparable to the colour changes observed after reaction of the neat ionic liquids with liquid mercury (Figure 5). As a useful laboratory application, Figure 9 shows a bed of chlorocuprate(II)-containing SILP used as a mercury trap on the off-gas line from a PSA Sir Galahad mercury analyser in our laboratory. The mercury breakthrough front can be clearly seen at the sharp transition from spent (pale green) absorbent to active (yellow-orange) material.

Interestingly, Haumann and co-workers have also recently reported the use of chlorocuprate(II)-containing SILPs for gas absorption applications. They investigated ammonia capture from air taking advantage of the strong coordination of nucleophilic amines to copper, and suggested their potential uses in broadband filters for breathing apparatus.

**Copper(II) supported ionic liquids at scale**

The laboratory results demonstrated that efficient and comprehensive scrubbing of mercury vapour from gas streams could be achieved with excellent capture kinetics. The accelerated breakthrough screening used residence times in the order of 50 ms in the test rig, and breakthrough times correlated with the theoretical maximum mercury capacities based on the mechanisms detailed in Equations (1) and (2). Following on from these small-scale laboratory tests, larger batches (ca. 100 g) of chlorocuprate(II) SILPs were prepared and tested in a 200 cm³ pilot-scale slipstream scrubber on-site at a Gas Processing Plant in Malaysia. The pilot trials (over 60 days) yielded a gas stream with mercury outlet concentrations of <0.01 µg m⁻³ (Figure 10), at least one order of magnitude below the sales specification for the natural gas.

This allowed a smooth transition to full plant implementation using 30 m³ of a SILP (see Figure 11). After three years of continuous operation, the outlet mercury concentration from the plant remained low, meeting the plant outlet specifications. This represents a remarkable success of transferring laboratory chemistry to full scale plant operation in a remarkably fast implementation (less than four years). The SILP is now commercialised via a marketing licensing agreement between PETRONAS and Clariant.
Conclusions

We have demonstrated that it is possible to develop in the laboratory, processes with industrial applicability. Starting at the gram scale, the translation of the ionic liquid chemistry into a SILP material that was scaled to full industrial implementation within four years has been exemplified. The fundamental chemistry (namely oxidative dissolution and complexation of mercury into the ionic liquid as a stable anionic mercurate(II) species) does not change with the scale of the process and the ability to digest large quantities of mercury (up to 20 wt%), coupled with efficient gas-contacting, made possible by formation of a SILP, underpins the successful development of these materials as commercially competitive solid phase mercury adsorbents in the natural gas industry.

Acknowledgments

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Notes and references

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