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Development of a PtSn bimetallic catalyst for direct fuel cells using bio-butanol fuel†

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Pt and PtSn catalysts were studied for n-butanol electro-oxidation at various temperatures. PtSn showed a higher activity towards butanol electro-oxidation compared to Pt in acidic media. The onset potential for n-butanol oxidation on PtSn is ~520 mV lower than that found on Pt, and significantly lower activation energy was found for PtSn compared with that for Pt.

Low molecular weight alcohols, such as methanol and ethanol, have been proposed as promising alternative fuels to H2 in low-temperature fuel cells due to their liquid nature, high energy density, low toxicity, availability and ease of handling. Methanol and ethanol are widely studied as fuels for direct alcohol fuel cells (DAFCs). Methanol is toxic, whereas ethanol is non-toxic and can be produced by fermentation of sugar. However, there is an on-going public debate on the ethics of utilizing food stocks for fuel production instead of nutrition. An alternate renewable source, which does not compete with food, is required for sustainable long-term energy production. Thus, second generation bio-fuels are proposed to be produced from non-food based biomass feedstock such as lignocellulose biomass (LCB). n-Butanol (n-BtOH) is considered as a 2nd generation bio-fuel, with better infrastructure compatibility than ethanol; it has higher energy density, lower water adsorption, and better blending ability with gasoline than ethanol and is considered as a potential substitute for gasoline in automobile fuel. Also, bio-butanol is non-poisonous, non-corrosive, biodegradable and does not lead to soil and water pollution. Multiple companies are working on the development of commercial scale bio-butanol production. For example, GreenBiologics UK, and Cobalt Technologies USA are involved in bio-butanol production. Butamax, a joint venture between BP and DuPont, has developed technology to produce bio-butanol from food stock biomass and in the second phase, is planning to produce butanol from LCB. Syntec Biofuels, Canada, uses a thermo-chemical method to produce bio-butanol from waste biomass (municipal solid waste, agricultural and forestry wastes).

Although there are a significant number of studies examining small-chain alcohols, such as methanol and ethanol, as potential fuels for fuel cells, there are fewer studies on longer-chain alcohols such as butanol. On complete electro-oxidation, butanol can provide 24e− per molecule compared to 12e− and 6e− for ethanol and methanol, respectively (eqn (1) for the general electro-oxidation of mono alcohols). However, increasing the number of carbon chain will increase the number of possible intermediates and complicate the reaction pathway.

\[
C_nH_{2n+1}OH + (2n - 1)H_2O \rightarrow nCO_2 + 6nH^+ + 6ne^- \tag{1}
\]

Butanol also has an advantage over shorter chain length alcohols in that the fuel cross-over through the electrolyte membrane, which is one of the major issues in DAFCs, will be reduced. Takky et al. studied butanol electro-oxidation on different electrodes (Pt, Au, Rh, Pd) in alkaline media. They have observed that n-butanol and iso-butanol show similar activity, whereas 2-butanol was less reactive. In contrast, tert-butanol was not oxidizable at all at room temperature. Only n-butanol and iso-butanol are produced from biomass feedstocks. Of the two, n-butanol appears to be the most promising as a potential fuel for DAFCs. In the work reported in this communication, the electro-oxidation of n-butanol is explored at both Pt and bimetallic PtSn electrocatalysts, building upon the previously published studies, which only described the monometallic catalyst.

The electrochemical measurements were run using a home built water-jacket three-electrode cell at controlled temperatures using a Bio-Logic EC Lab SP-200 potentiostat connected to a PC with EC-Lab software. A Ag/AgCl in 3 M NaCl (BASi, USA)
(0.210 V vs. SHE at 25 °C) reference electrode was used, and a Pt mesh (Goodfellow, UK) attached to a Pt wire (Mateck, Germany) was used as counter electrode. All potentials are reported with respect to Ag/AgCl - 3 M NaCl unless otherwise specified. Pt was electro-deposited on a glassy carbon (GC) substrate (area = 0.38 cm²) from 5 mM H₂PtCl₆ + 0.1 M HCl aqueous solution and Sn was electro-deposited on the Pt from 1 mM SnCl₂ + 0.5 M H₂SO₄ solution at a deposition potential of −0.210 V. The Sn coverage can be controlled by Sn deposition time; for example, the 25% Sn coverage can be achieved in 20 seconds. The electro-deposited Pt has ‘cauliflower-like’ spherical morphology and no significant morphology change occurred upon the Sn deposition on the Pt from the SEM analysis (ESI,† Fig. S1 and S2). The cyclic voltammogram (CV) of Pt showed characteristic features of polycrystalline Pt in acidic solution (ESI,† Fig. S3) and the Pt active area (Aᵣ) was calculated using eqn (2).

\[
Aᵣ = \frac{Q_{H}}{Q_{H}^o}
\]

where \(Q_{H}\) is the charge for hydrogen desorption and \(Q_{H}^o\) is the charge required to oxidize a monolayer of hydrogen from Pt surface (0.21 mC cm⁻²). The values of current were divided by \(Aᵣ\) to provide the specific current density rather than the geometric current density so as to provide a better comparison of the true catalytic activity per surface site. Fig. 1 shows the CV of Pt in n-butanol solution. The CV features of Pt in supporting acidic solution at the hydrogen adsorption/desorption potential region were inhibited in the CV of butanol containing solution indicating that some of the Pt surface sites were covered by the adsorbates of butanol. The current was negligible at lower potential during the positive going potential sweep (PGPS) and started to increase at potentials > 0.55 V. Two oxidation peaks were observed during the PGPS at ~0.73 V (peak a₁) and ~1.07 V (peak a₂). During the negative going potential sweep (NGPS), the current was positive till 0.70 V and below which a small cathodic peak appeared (peak c₁) followed by an oxidation peak at an onset potential of ~0.52 V and peak maximum at ~0.38 V (peak a₃).

It is noted that the onset of n-butanol oxidation, peak a₁, intersects with that of the water oxidation on Pt (ESI,† Fig. S3) indicating that the formation of the surface OH ads may facilitate n-butanol oxidation on Pt. Increasing the potential promotes the formation of surface OH ads which in turn enhances the n-butanol oxidation (peak a₂). However, at potentials > 1.0 V the Pt surface is covered by the oxide species leading to a decrease in n-butanol oxidation current. The reduction of the oxide species during the NGPS allows fresh adsorption and oxidation of butanol giving peak a₃.

At higher temperatures, a significant improvement in n-butanol oxidation was observed. An increase in current with a negative shift in the onset potential and a positive shift in the re-oxidation peak was observed with the increase in temperature (Fig. 2). The activation energy (\(E_a\)) for the pure Pt in the n-butanol oxidation calculated from the Arrhenius plot (ln j vs. 1/T) was ~25 kJ mol⁻¹ and ~15 kJ mol⁻¹ at 0.70 V and 1.00 V, respectively (Fig. 3a). A higher activation energy for the first anodic peak compared to the second anodic peak was reported in literature for alcohol oxidation reaction on Pt in acidic media.⁴⁻¹³⁻¹⁴

It is well-documented that addition of oxophilic metals such as Ru and Sn to Pt can promote surface oxidant OH ads formation at lower potential. This is expected to enhance the alcohol oxidation through the bi-functional mechanism. PtRu is known to be the
A PtSn electrode was prepared by electrochemical deposition of Sn on Pt/GC. The CVs of Pt and PtSn electrodes in 0.1 M H₂SO₄ supporting electrolyte and n-butanol containing solutions are compared in Fig. 4. The upper potential in CV analysis was limited to 0.75 V in the case of the PtSn electrode in order to avoid any electro-dissolution of Sn at higher potentials. The hydrogen adsorption/desorption region was suppressed on Sn addition due to the blockage of the Pt sites by Sn for hydrogen adsorption. The Sn coverage (θSn) on Pt was calculated from the change in hydrogen desorption charge before and after the Sn deposition on Pt/GC according to eqn (3), where QH is the hydrogen desorption charge for the PtSn electrode and QH is the hydrogen desorption charge for Pt. The θSn calculated from the CV for the PtSn was ~25%.

For the PtSn in H₂SO₄ solution (see Fig. 4a), surface oxidation appears with an onset potential of ~0.10 V and peak potential at around 0.55 V indicating the oxidation of the adsorbed Sn. The hydrogen adsorption/desorption features are preserved for the PtSn electrodes but with reduced area.

Fig. 4b compares the PGPS voltammograms of the Pt and PtSn electrodes in 0.1 M n-BtOH + 0.1 M H₂SO₄. The Pt shows the characteristic features in acid with a n-BtOH oxidation onset potential of ~0.55 V. A similar onset potential was reported in literature for n-butanol oxidation on Pt in acidic media. However, for the PtSn electrode, a pre-peak at ~0.27 V (onset potential at ~0.03 V) was observed before the a₁ peak on Pt. This peak was only observed in the n-butanol containing solution, indicating that the peak originated from butanol oxidation. This was further confirmed by adding butanol gradually to the sulphuric acid supporting electrolyte while running the CV and the pre-peak appears only when butanol is added to the acidic solution. In fact, Sn coverage of 10–30% was found to be optimum (ESI, Fig. S4). At higher Sn coverage (~40%), the butanol oxidation peak a₁ was found to be suppressed significantly (ESI, Fig. S5). Since this pre-peak appears in a potential region far below PtOH/PtO₂ formation, the active oxygen species for the n-butanol oxidation must originate from the Sn and is probably due to the low onset potential for OH formation on Sn observed previously.

$$\theta_{Sn} = \frac{Q_H - Q_{Sn}}{Q_H}$$  \hspace{1cm} (3)

It is to be noted that onset potential for n-butanol oxidation on PtSn is significantly lower than that at Pt (~520 mV lower). Also the current density of the pre-peak was ~2–3 times higher than peak a₁ on pure Pt showing the addition of Sn facilitates the oxidation reaction. The E_b calculated at 0.25 V for the PtSn was ~11 kJ mol⁻¹ (Fig. 3b, ESI, Fig. S3) which is significantly lower than that at peak a₁ on pure Pt (25 kJ mol⁻¹). Taky et al. reported an E_b of 11.7 and 14.2 kJ mol⁻¹ for n-butanol oxidation on Pt in alkaline media at potentials of 0.39 V and 0.29 V, respectively. It should be noted that in alkaline medium the Pt could be covered by OHads even at low potentials and thus resembles our PtSn system. The lower E_b on PtSn at lower potential may be attributed to an increase in the effectiveness of the dissociative adsorption and oxidation of n-butanol at the PtSn catalyst electrode.

In order to have an understanding of the oxidation peak behaviour with other alcohol molecules, PtSn for ethanol (EtOH) and iso-butanol (iso-BtOH) (ESI, Fig. S6) were compared with that of n-butanol. The pre-peak was also present in both iso-BtOH and EtOH solution. It is interesting to note that even though the onset and half wave potentials of ethanol oxidation (~0.68 V and ~0.49 V, respectively) and n-butanol oxidation (~0.55 V and ~0.64 V) were different on pure Pt, the onset and half wave potentials for all alcohols were the same for PtSn (~0.02 V and ~0.16–0.20 V respectively) (Table 1). This result is significant as it indicates that on PtSn the oxidation is initiated and determined by the water oxidation on Sn, rather than on the nature of the molecule. González et al. observed that alcohols with H atoms on the β carbon can be oxidized on PtSn electrodes at lower potential than on monometallic Pt electrode. They proposed that H abstraction from α and β carbon could lead to the formation of a stabilizing enol structure, which could convert to aldehyde. This conversion would be facilitated by the OH/O-rich PtSn surface as the attack on enol by adsorbed H₂O could easily form a hydrated aldehyde.

In conclusion, significantly higher activity of a PtSn bimetallic catalyst for n-butanol electrode-oxidation in acidic media compared to pure Pt has been clearly demonstrated. On addition of Sn, a distinguishable oxidation pre-peak was observed at a much lower onset potential for PtSn compared to those for Pt. This has been attributed to facile formation of OHads at lower potentials on Sn. These results suggest that PtSn could be an effective catalyst for n-butanol electro-oxidation in direct alcohol fuel cells.

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**Table 1** The onset potential and half wave potential/E_b (in parenthesis) of Pt and PtSn in 0.1 M alcohol + 0.1 M H₂SO₄ solution at room temperature.

<table>
<thead>
<tr>
<th>Catalyst</th>
<th>Butanol</th>
<th>iso-Butanol</th>
<th>Ethanol</th>
</tr>
</thead>
<tbody>
<tr>
<td>Pt</td>
<td>~0.55</td>
<td>~0.55</td>
<td>~0.40</td>
</tr>
<tr>
<td></td>
<td>(~0.64)</td>
<td>(~0.62)</td>
<td>(~0.56)</td>
</tr>
<tr>
<td>PtSn</td>
<td>~0.03</td>
<td>~0.03</td>
<td>~0.03</td>
</tr>
<tr>
<td></td>
<td>(~0.16)</td>
<td>(~0.16)</td>
<td>(~0.20)</td>
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Notes and references


